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[MICROSOFT WORD - COMBINED GAS LAWS-NOTES](#)

Combined Gas Laws PV = constant PV = constant (Boyles Law) T (combined gas law) V = constant (Charles Law) T P = constant TALL gas law problems must be work using absolute temperature, kelvin (K).Memory Work: 760 mmHg = 760 torr = 1 atm = 1.013 X 10⁵ Pa and P₁V₁ = P₂V₂ T₁ T₂ or P₁V₁T₂ = P₂V₂T₁ and STP (standard temperature and pressure) 0 oC and 1 atm (273 K) (760 torr)V Volume amount of gas cons...

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[FORM 94 - SEARCH WARRANT \(NATIONAL ELECTRICITY LAW S.21 OR NATIONAL GAS LAW S.35\)](#)

Form 94 SEARCH WARRANT Magistrates Court of South Australia www.courts.sa.gov.au National Electricity Law section 21 or National Gas Law section 35Registry: .. Court File No:..Authorised person under section 20 (Electricity) or 32 (Gas):Name: , ID No: Surname G...

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[GAS LAW REFERENCE SHEET](#)

Dalton s Law of Partial Pressures P_{total} = P₁ + P₂ + P₃ . . . Charles Law Boyle s Law V₁ = V₂ V₁P₁ = V₂P₂ T₁ T₂ Gay-Lussacs Law P₁ = P₂ T₁ T₂ Combined Gas Law Ideal Gas Law V₁P₁ = V₂P₂ PV = nRT T₁ T₂ Some Pressure Conversion Factors1 atmosphere = 760 mm Hg = 760 torr = 101.325 kPa Values for Standard Temperature and Pressure Pressure = 760 mm Hg Temperature = 273 K R Values .0821 atm L 62.4 mm ...

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[EXPT 07 - THE IDEAL GAS LAW AND DENSITIES](#)

EXPERIMENT 7THE IDEAL GAS LAW AND DENSITYIn this experiment you will determine the average molecular mass of air using two differentmethods first by measuring the density of air with the density of several different gases andsecondly by using the Ideal Gas LawTo determine density of a gas you will fill a syringe with a gas and find it s mass The volume ofthe gas will be determined by weighing a eq...

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