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[WORKSHEET - BALANCING CHEMICAL EQUATIONS](#)

Worksheet: Writing and Balancing Chemical Reactions
 1. Balance the following equations and indicate the type of reaction as formation, decomposition, single replacement, double replacement, hydrocarbon combustion, or other.
 a. $\text{Cu (s)} + \text{O}_2 \text{ (g)} \rightarrow \text{CuO (s)}$
 b. $\text{H}_2\text{O (l)} \rightarrow \text{H}_2 \text{ (g)} + \text{O}_2 \text{ (g)}$
 c. $\text{Fe (s)} + \text{H}_2\text{O (g)} \rightarrow \text{H}_2 \text{ (g)} + \text{Fe}_3\text{O}_4 \text{ (s)}$
 d. $\text{AsCl}_3 \text{ (aq)} + \text{H}_2\text{S (aq)} \rightarrow \text{As}_2\text{S}_3 \text{ (s)} + \text{HCl (aq)}$
 e. $\text{CuSO}_4 \cdot 5 \text{ H}_2\text{O (s)} \rightarrow \text{CuSO}_4 \text{ (s)} + \text{H}_2\text{O (l)}$

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[BALANCING CHEMICAL EQUATIONS](#)

BALANCING CHEMICAL EQUATIONS
 The Law of Conservation of Mass says that in a chemical reaction, matter is neither created nor destroyed. Practically speaking, whatever atoms begin in the reactants of a chemical reaction, the same number of atoms of each element must appear in the products. The process of balancing chemical equations is intended to reflect reality, that is, to give us a true picture of...

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Balancing Chemical Equations Answer Key
 Balance the equations below:
 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$
 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$
 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$
 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$
 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$
 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$
 7) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$
 8) $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$
 9) $2 \text{ C}_8\text{H}_{18} + 25 \text{ O}_2 \rightarrow 16 \text{ CO}_2 + 18 \text{ H}_2\text{O}$
 10) $1 \text{ FeCl}_3 + 3 \text{ NaOH} \rightarrow 1 \text{ Fe(OH)}_3 + 3 \text{ NaCl}$
 11) $4 \text{ P} + 5 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_5$
 12) $2 \text{ Na} + \dots$

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[CHEMICAL EQUATIONS \(13K\)](#)

Chemical equations Write balanced symbol equations for the following chemical reactions. Remember to balance the equations, and to include the state symbols.
 1. barium chloride + sodium sulphate \rightarrow barium sulphate + sodium chloride
 2. calcium + oxygen \rightarrow calcium oxide
 3. calcium carbonate + sulphuric acid \rightarrow calcium sulphate + carbon dioxide + water
 $\text{H}_2\text{SO}_4 \text{ (aq)}$
 4. carbon + oxygen \rightarrow carbon dioxide
 5. copper...

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